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Reaction Rates And Equilibrium Worksheet Answers Chapter 18

Section 18.2
Reversible Reactions and Equilibrium

OBJECTIVES

- Describe how the amounts of reactants and products change in a chemical system at equilibrium.
- Identify three stresses that can change the equilibrium position of a chemical system.
- Explain what the value of K_{eq} indicates about the position of equilibrium.

Reversible Reactions

- Some reactions do not go to completion as we have assumed.
 - They may be _____ a reaction in which the conversion of reactants to products and the conversion of products to reactants occur simultaneously.
 - $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$
 - $2SO_3(g) \rightleftharpoons 2SO_2(g) + O_2(g)$
- The two equations can be combined into one, by using a _____, which tells us that it is a reversible reaction:
 $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$
 - A chemical equilibrium occurs, and _____ change occurs in the actual amounts of the components of the system.
- Even though the rates of the forward and reverse are equal, the _____ of components on both sides may not be equal.
 - An equilibrium position may be shown:
A: _____ B: _____ or A: _____ B: _____
1% 99% 99% 1%
 - Note the emphasis of the arrows direction.
 - It depends on which side is favored; almost all reactions are reversible to some extent.

Le Chatelier's Principle

- The French chemist Henri Le Chatelier (1850-1936) studied how the equilibrium position shifts as a result of changing conditions.
- Le Chatelier's principle: _____

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